

Exchange current density and Tafel plot

Butler volmer equation:

$$i = i_c - i_a = i_0 \left[ \exp\left(\frac{\alpha z F \eta}{RT}\right) - \exp\left(\frac{-(1-\alpha) z F \eta}{RT}\right) \right] \quad \text{(A)}$$

If the current density  $i$  is plotted against the overpotential in accordance with equ. (A) Two cases are shown:

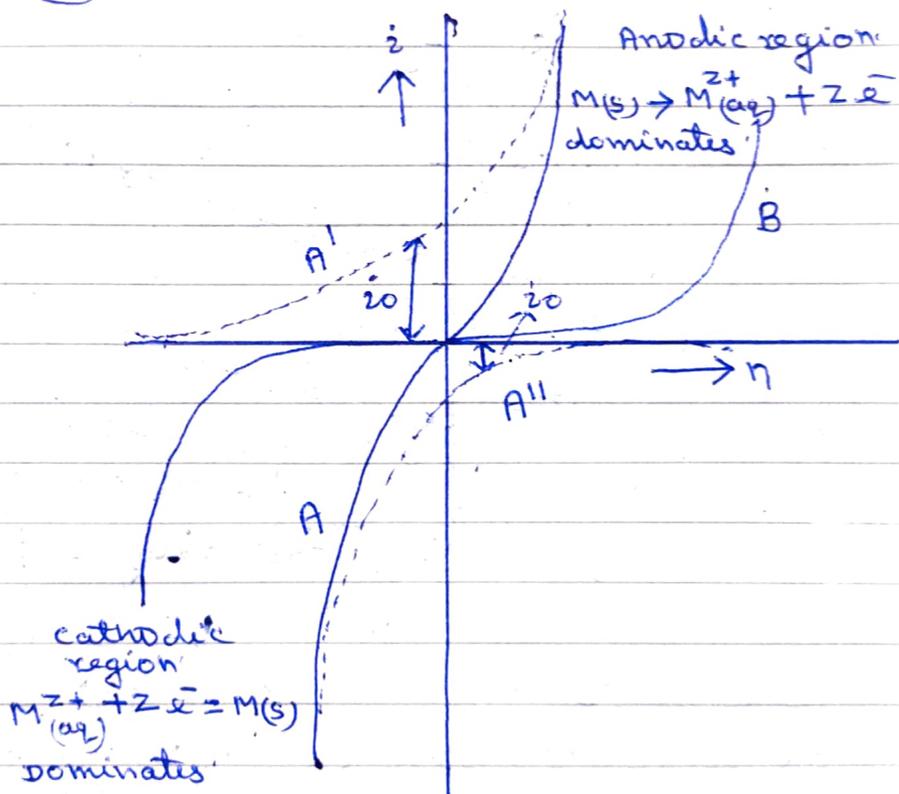


Fig (1)

one labelled 'A' is a case in which there is rather high exchange current  $i_0$  at both electrodes. The other case labelled 'B' corresponds to a very low exchange current  $i_0$ . In this case, a large overpotential is required to cause appreciable current

How the cell

We see that, factor that determines the activation overpotential, in accord with Butler-Volmer equation, is the exchange current density  $i_0$ . The transfer factor  $\alpha$  influences the shape of the current density  $i$  vs overpotential  $\eta$  curve as fig II

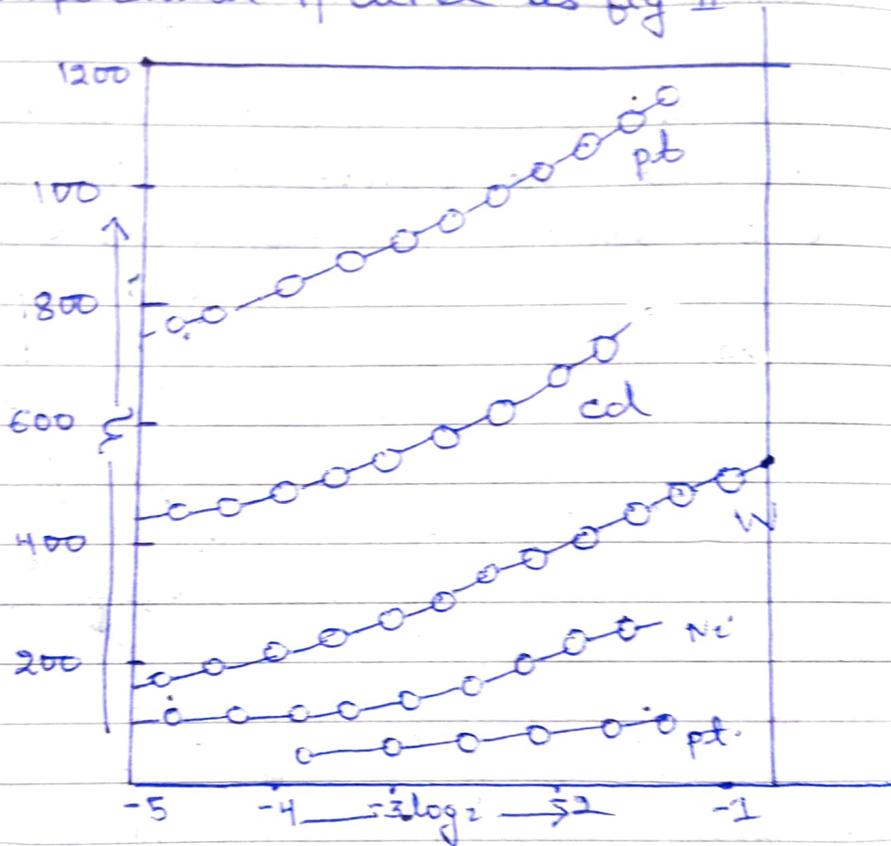


fig. II

one way to measure  $\alpha$  to fit the experimental curve to the Butler-Volmer equ.

Tafel plot or Equations :-

If the overpotential has large +ve or (-ve) values  $\eta \gg RT/nF$ , one of the partial currents becomes much greater than the other, which is negligible in this case.

either

$$\ln i_a = \ln i_0 - [(1-\alpha)ZF/RT]n \quad \text{--- (B)}$$

$$\text{or } \ln i_c = \ln i_0 + (\alpha ZF/RT)n \quad \text{--- (C)}$$

This type of logarithmic dependence of  $i$  on  $n$  was found ~~empirically~~ empirically in 1905 by Tafel.

The slope of linear  $\ln i$  vs  $n$  plots gives the transfer coefficient  $\alpha$  and the intercept gives the exchange current density  $i_0$  by fig. II.

Over potential can be caused by slow reaction in the solution adjacent to an electrode, and in the process of deposition of a solid product on an electrode. In many cases, a diffusion overpotential occurs together with an activation over potential and methods are available for separating these factor. Thus, a great variety of interesting chemical kinetics can be studied by electronic techniques. practical application of this branch are found in such fields as fuel cells, storage batteries for vehicle, electro-chemical synthesis. etc.